

WEST Search History

DATE: Friday, September 26, 2003

<u>Set Name</u>	<u>Query</u>	<u>Hit Count</u>	<u>Set Name</u>
side by side			result set
<i>DB=USPT,PGPB,JPAB,EPAB,DWPI,TDBD; PLUR=YES; OP=ADJ</i>			
L11	l9 and L10	160	L11
L10	density or dense or densify	1160908	L10
L9	l6 and L8	367	L9
L8	pressur\$8 same temperature	732414	L8
L7	pressur\$8 and L6	424	L7
L6	l1 same l2 same l3	578	L6
L5	pressur\$8 and L4	1054	L5
L4	l1 and l2 and L3	1495	L4
L3	ru or ruthenium or sruuo?sub.3 or ruthenate	672622	L3
L2	sr or strontium or sruuo?sub.3	171916	L2
L1	perovskit\$4	13958	L1

END OF SEARCH HISTORY

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(FILE 'HOME' ENTERED AT 10:30:03 ON 26 SEP 2003)

FILE 'CAPLUS' ENTERED AT 10:30:24 ON 26 SEP 2003

L1 29555 S PEROVSKIT?
L2 46165 S PRESS? WITH TEMP?
L3 358 S L1 AND L2
L4 431180 S DENS?
L5 35 S L3 AND L4
L6 54 S DENS? WITH PEROVSKIT?
L7 323 S L3 NOT L5
L8 323 S L7 NOT L6

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(FILE 'HOME' ENTERED AT 09:37:18 ON 26 SEP 2003)

FILE 'CAPLUS' ENTERED AT 09:38:29 ON 26 SEP 2003

L1 29555 S PEROVSKIT?
L2 418308 S PRESS? AND TEMP?
L3 46165 S PRESS? WITH TEMP?
L4 358 S L3 AND L1
L5 431180 S DENS?
L6 35 S L4 AND L5
L7 54 S DENS? WITH PEROVSKIT?

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L6 ANSWER 1 OF 35 CAPLUS COPYRIGHT 2003 ACS on STN

AN 2003:133753 CAPLUS

DN 138:157627

TI High pressure sintering of ruthenium-based **perovskites** for improved **densification** and minimized crystal growth

IN Vagarali, Suresh S.; Lucek, John W.

PA USA

SO U.S. Pat. Appl. Publ., 4 pp.

CODEN: USXXCO

DT Patent

LA English

IC ICM C04B033-32

ICS C04B035-01

NCL 264681000; 423420200; 423593000

CC 57-2 (Ceramics)

Section cross-reference(s): 56, 77

FAN.CNT 1

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
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PI US 2003034591 A1 20030220 US 2001-931312 20010816

PRAI US 2001-931312 20010816

AB A **perovskite** feedstock (such as RuSrO₃ powders or preforms) is placed in a high-pressure cell (with graphite or NaCl pressure-transmitting media) of a high **pressure/high temp.** app. and subjected to pressures .gtorsim.2 kbar and temps. .gtorsim.800.degree.. for a time adequate to increase the d. of the preform. The sintering under pressure allows for lower temps. and minimizes grain growth.

ST pressure sintering **density** ruthenium strontium **perovskite** crystal growth

IT **Perovskite**-type crystals

(d. of; high pressure sintering of ruthenium-based **perovskites** for improved **densification** and minimized crystal growth)

IT Apparatus

(high-**pressure/high-temp.**; high pressure sintering of ruthenium-based **perovskites** for improved **densification** and minimized crystal growth)

IT Sintering

(hot pressing; high pressure sintering of ruthenium-based **perovskites** for improved **densification** and minimized crystal growth)

IT Crystal growth

(**perovskite**-type; high pressure sintering of ruthenium-based **perovskites** for improved **densification** and minimized crystal growth)

IT Crystallization

(pressure-induced; high pressure sintering of ruthenium-based **perovskites** for improved **densification** and minimized crystal growth)

IT 7439-98-7, Molybdenum, uses

RL: NUU (Other use, unclassified); USES (Uses)

(foil; high pressure sintering of ruthenium-based **perovskites** for improved **densification** and minimized crystal growth)

IT 12169-14-1, Ruthenium strontium oxide (RuSrO₃)

RL: PEP (Physical, engineering or chemical process); PRP (Properties); PYP (Physical process); PROC (Process)

(**perovskite**, ferromagnets; high pressure sintering of ruthenium-based **perovskites** for improved

densification and minimized crystal growth)

IT 7647-14-5, Sodium chloride, uses 7782-42-5, Graphite, uses

RL: NUU (Other use, unclassified); USES (Uses)

(pressure-transmitting media; high pressure sintering of

Applicants

ruthenium-based **perovskites** for improved
densification and minimized crystal growth)

L6 ANSWER 14 OF 35 CAPLUS COPYRIGHT 2003 ACS on STN
AN 1998:540829 CAPLUS
DN 129:278608
TI Elasticity of CaTiO₃-CaSiO₃ **perovskites**
AU Sinelnikov, Yegor D.; Chen, Ganglin; Liebermann, Robert C.
CS Center of High Pressure Research and Department of Earth and Space Sciences, State University of New York at Stony Brook, Stony Brook, NY, 11794, USA
SO Physics and Chemistry of Minerals (1998), 25(7), 515-521
CODEN: PCMIDU; ISSN: 0342-1791
PB Springer-Verlag
DT Journal
LA English
CC 53-3 (Mineralogical and Geological Chemistry)
AB Polycryst. specimens in the CaTiO₃-CaSiO₃ **perovskite** system have been hot-pressed in a 2000-ton uniaxial split-sphere app. (USSA-2000) at pressures up to 15 GPa and temp. of 1550.degree.C, for the compns. CaTiO₃, Ca(Ti0.75Si0.25)O₃, Ca(Ti0.5Si0.5)O₃. For the specimens with the bulk **densities** within 1% of the X-ray d., compressional and shear wave velocity measurements have been conducted using ultrasonic interferometry. The measured adiabatic bulk moduli (K_s) for the CaTiO₃ and Ca(Ti0.5Si0.5)O₃ **perovskites** are 175(1) and 188(1) GPa and shear moduli (G) of 106(1) and 109(1) GPa. In situ X-ray diffraction studies at high **pressure** and **temp.** resulted in isothermal values for K₀ of 170(5) and 185(5) GPa, resp. For the unquenchable CaSiO₃ **perovskite**, elasticity theory and systematics were used to predict K₀=212(7) GPa and G₀=112(5) GPa; this shear modulus is 37% less than that for (Mg,Fe)SiO₃ **perovskite**, suggesting that CaSiO₃ **perovskite** cannot be ignored in modeling the compn. of the Earth's lower mantle.
ST lower mantle calcium silicate titanate elasticity; **perovskite** silicate elasticity lower mantle model
IT Elasticity
Simulation and Modeling, physicochemical
Young's modulus
(lower mantle-modeling implications of the elastic modulus for phases of the **perovskite**-structured solid-soln. series CaTiO₃-CaSiO₃)
IT Mantle (earth)
(lower mantle; lower mantle-modeling implications of the elastic modulus for phases of the **perovskite**-structured solid-soln. series CaTiO₃-CaSiO₃)
IT **Perovskite**-type crystals
(of CaTiO₃-CaSiO₃ solid-soln. series; lower mantle-modeling implications of the elastic modulus for phases of the **perovskite**-structured solid-soln. series CaTiO₃-CaSiO₃)
IT 12049-50-2, Calcium titanium oxide (CaTiO₃)
RL: GPR (Geological or astronomical process); PRP (Properties); PROC (Process)
(lower mantle-modeling implications of the elastic modulus for phases of the **perovskite**-structured solid-soln. series CaTiO₃-CaSiO₃)
IT 10101-39-0, Silicic acid (H₂SiO₃), calcium salt (1:1) 12194-71-7,
Perovskite 192707-00-9, Calcium titanium silicate (Ca₂TiSiO₆)
198972-70-2, Calcium titanium oxide silicate (CaTi0.75O₂(SiO₄)_{0.25})
RL: GPR (Geological or astronomical process); PRP (Properties); PROC (Process)
(**perovskite**-structured, elastic modulus of; lower mantle-modeling implications of the elastic modulus for phases of the **perovskite**-structured solid-soln. series CaTiO₃-CaSiO₃)
RE.CNT 42 THERE ARE 42 CITED REFERENCES AVAILABLE FOR THIS RECORD
RE

L6 ANSWER 24 OF 35 CAPLUS COPYRIGHT 2003 ACS on STN
AN 1992:536222 CAPLUS
DN 117:136222
TI Preparation of high-**density perovskite**-type lanthanum nickel trioxide by using hot isostatic pressing (HIP)
AU Kamata, Kiichiro; Ohshio, Shigeo; Tanaka, Kiyoshi; Kawai, Koh; Yamada, Tomoo; Nishino, Junichi
CS Dep. Chem., Nagaoka Univ. Technol., Nagaoka, 940-21, Japan
SO Nippon Seramikkusu Kyokai Gakujutsu Ronbunshi (1992), 100(July), 972-4
CODEN: NSKRE2; ISSN: 0914-5400
DT Journal
LA Japanese
CC 57-2 (Ceramics)
Section cross-reference(s): 76
AB Dense sintered bodies of **perovskite**-type LaNiO₃, which shows metallic conduction, were prepd. by HIP without flux such as Na₂CO₃. A stoichiometric mixt. of La₂O₃ and NiO was precalcined in air at 900.degree. for 5 h. The ground sample was then sintered at 1100 or 1300.degree. in air for 20 h. The sintered bodies were composed of La₂NiO₄ and NiO and had d. .ltoreq.95%. Single-phase **perovskite**-type LaNiO₃ of d. 95% was obtained from its presintered body by HIP under 101 MPa (in Ar gas contg. 20% O) at 1400.degree. for 2 h. A very low resistivity of 2.2 .times. 10⁻⁶ .OMEGA. m was measured for this hot-isostatically pressed body of LaNiO₃. The stability region for the **perovskite**-type LaNiO₃ in the **pressure-temp.** diagram is discussed.
ST lanthanum nickel oxide **perovskite** prep; hot isostatic pressing
lanthanum nickel oxide
IT Ceramic materials and wares
(lanthanum nickel oxide, prep. of, by hot isostatic pressing)
IT Electric resistance
(of lanthanum nickel oxide ceramics, prep. condition effect on)
IT Sintering
(hot isostatic pressing, of lanthanum nickel oxide, phase compn. and elec. resistance in relation to)
IT 12031-18-4P, Lanthanum nickel oxide (LaNiO₃)
RL: SPN (Synthetic preparation); PREP (Preparation)
(**perovskite**-type, ceramics, prep. of, by hot isostatic pressing)

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L7 ANSWER 30 OF 54 CAPLUS COPYRIGHT 2003 ACS on STN
AN 1996:93261 CAPLUS
DN 124:183397
TI Enhanced **densification** of **perovskite**-type strontium titanate ceramics
AU Liu, Kuo-Shung; Lin, I-Nan
CS Department Materials Science and Engineering, National Tsing-Hua University, Hsinchu, 30043, Taiwan
SO ISAF '94, Proceedings of the IEEE International Symposium on Applications of Ferroelectrics, 9th, University Park, Pa., Aug. 7-10, 1994 (1994), 261-4. Editor(s): Pandey, R. K.; Liu, Michael; Safari, Ahmad. Publisher: Institute of Electrical and Electronics Engineers, New York, N. Y.
CODEN: 62GYAM
DT Conference
LA English
CC 57-2. (Ceramics)
AB The densification of SrTiO₃ materials was enhanced by novel sintering method. On the other hand, hot isostatic pressing and microwave sintering may markedly enhance diffusion of ions such that the undesirable features of mixed oxide powders are easily overcome and the materials well sintered. Densification of the materials has already started when microwave-sintering at 1220.degree., but crystal growth occurs only at 1250.degree.. Migration of grain boundaries can be effectively triggered by applying a short temp. pulse at the onset of the densification process.
ST sintering strontium titanate dielec ceramic; titania strontium carbonate sintering ceramic; ferroelec perovskite crystal ceramic
IT Perovskite-type crystals
(enhanced **densification** of **perovskite**-type strontium titanate ceramics)
IT Ferroelectric substances
(ceramic, enhanced **densification** of **perovskite**-type strontium titanate ceramics)
IT Sintering
(hot isostatic pressing, enhanced **densification** of **perovskite**-type strontium titanate ceramics)
IT Sintering
(microwave, enhanced **densification** of **perovskite**-type strontium titanate ceramics).
IT 1633-05-2, Strontium carbonate 12060-59-2, Strontium titanate 13463-67-7, Titania, processes
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(enhanced **densification** of **perovskite**-type strontium titanate ceramics)

8 ANSWER 46 OF 323 CAPLUS COPYRIGHT 2003 ACS on STN
AN 2001:99813 CAPLUS
DN 134:289360
TI Valence variations in titanium-based **perovskite** oxides by high-pressure and high-temperature method
AU Li, Liping; Li, Guangshe; Miao, Jipeng; Su, Wenhui; Inomata, Hiroshi
CS Department of Physics, Jilin University, Changchun, 130023, Peop. Rep. China
SO Journal of Materials Research (2001), 16(2), 417-424
CODEN: JMREEE; ISSN: 0884-2914
PB Materials Research Society
DT Journal
LA English
CC 78-2 (Inorganic Chemicals and Reactions)
Section cross-reference(s): 73, 75
AB Typical Ti-based **perovskite** oxides $Eu_{1-x}Ba_xTiO_3$ ($x = 0.6-0.8$), $Eu_{1-x}K_xTiO_3$ ($x = 0.2, 0.32$), and $La_{0.7}(M)0.3TiO_3$ ($M = Na, K$) were synthesized by high **pressure** and **temp.** using RE_2O_3 ($RE = La, Eu$), TiO_2 , alk., or alk. earth carbonates as the starting materials. X-ray diffraction data anal. showed that there was a structural transformation in $Eu_{1-x}Ba_xTiO_3$ by varying Ba content [i.e., from cubic ($x = 0.6, 0.7$) to tetragonal ($x = 0.8$)], and that samples $Eu_{1-x}K_xTiO_3$ and $La_{0.7}(M)0.3TiO_3$ crystd. in the cubic **perovskite** structure. ^{151}Eu Mossbauer spectroscopy and ESR measurements revealed mixed valence of Eu^{2+}/Eu^{3+} in samples $Eu_{1-x}Ba_xTiO_3$ and $Eu_{1-x}K_xTiO_3$, while Ti ions were present in pure Ti^{4+} state. Cubic $Eu_{1-x}K_xTiO_3$ was metastable, which decompd. into a mixt. of **perovskite** and pyrochlore phases at high temps. as accompanied by an oxidn. process from Eu^{2+} to Eu^{3+} . For samples $La_{0.7}(M)0.3TiO_3$, Ti^{3+} signals were clearly obsd. The redn. mechanisms for Eu ions at A site and Ti ions at B site in the **perovskite** oxides are discussed in terms of the chem. nature of the framework ions and substitution ions under high **pressure** and **temp.**
ST mixed valence quaternary titanium oxide **perovskite** prepns;
pressure effect valence quaternary titanium oxide **perovskite**;
temp effect valence quaternary titanium oxide **perovskite**;
crystal structure quaternary titanium oxide **perovskite**;
Moessbauer europium 151 quaternary titanium oxide **perovskite**
IT Mossbauer effect
(europium-151; in barium europium titanium oxide **perovskites**)
IT Pressure
(high; valence variations in titanium-based **perovskite** oxides by high-pressure and high-temp. method)
IT Crystal structure
(of quaternary titanium oxide **perovskites** from XRD data)
IT Mixed valence
(valence variations in titanium-based **perovskite** oxides by high-pressure and high-temp. method)
IT 513-77-9, Barium carbonate
RL: RCT (Reactant); RACT (Reactant or reagent)
(for prepns. of quaternary barium europium titanium oxide **perovskites** by high-pressure and high-temp. method)
IT 1308-96-9, Europium oxide (Eu $2O_3$)
RL: RCT (Reactant); RACT (Reactant or reagent)
(for prepns. of quaternary europium barium/potassium titanium oxide **perovskite** by high-pressure and high-temp. method)
IT 1312-81-8, Lanthanum oxide (La $2O_3$)
RL: RCT (Reactant); RACT (Reactant or reagent)
(for prepns. of quaternary lithium/potassium lanthanum titanium oxide **perovskites** by high-pressure and high-temp. method)
IT 298-14-6
RL: RCT (Reactant); RACT (Reactant or reagent)

- (for prepn. of quaternary potassium rare earth titanium oxide **perovskite** by high-pressure and high-temp. method)
- IT 144-55-8, Sodium hydrogen carbonate, reactions
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (for prepn. of quaternary sodium lanthanum titanium oxide **perovskite** by high-pressure and high-temp. method)
- IT 13463-67-7, Titanium dioxide, reactions
 RL: RCT (Reactant); RACT (Reactant or reagent)
 (for prepn. of quaternary titanium-based oxide **perovskites** by high-pressure and high-temp. method)
- IT 272766-29-7P, Europium potassium titanium oxide (Eu0.8K0.2TiO₃)
 RL: PRP (Properties); SPN (Synthetic preparation); PREP (Preparation)
 (high-pressure and -temp. prepn., crystal structure
 lattice parameters, ESR, and europium-151 Moessbauer of)
- IT 152857-48-2P, Lanthanum sodium titanium oxide (La0.7Na0.3TiO₃)
 152857-52-8P, Lanthanum potassium titanium oxide (La0.7K0.3TiO₃)
 212179-79-8P, Barium europium titanium oxide (Ba0.8Eu0.2TiO₃)
 RL: PRP (Properties); SPN (Synthetic preparation); PREP (Preparation)
 (high-pressure and -temp. prepn., crystal structure
 lattice parameters, and ESR of)
- IT 212179-80-1P, Barium europium titanium oxide (Ba0.7Eu0.3TiO₃)
 212179-81-2P, Barium europium titanium oxide (Ba0.6Eu0.4TiO₃)
 RL: PRP (Properties); SPN (Synthetic preparation); PREP (Preparation)
 (high-pressure and -temp. prepn., crystal structure
 lattice parameters, and europium-151 Moessbauer of)
- IT 272766-30-0P, Europium potassium titanium oxide (Eu0.68K0.32TiO₃)
 RL: PRP (Properties); RCT (Reactant); SPN (Synthetic preparation); PREP (Preparation); RACT (Reactant or reagent)
 (high-pressure and -temp. prepn., crystal structure
 lattice parameters, thermal decomprn., ESR, and europium-151 Moessbauer of)
- IT 272766-30-0DP, Europium potassium titanium oxide (Eu0.68K0.32TiO₃),
 reaction products from sintering
 RL: PRP (Properties); SPN (Synthetic preparation); PREP (Preparation)
 (prepn., XRD, and europium-151 Moessbauer of **perovskite** and pyrochlore phases)
- IT 14378-48-4, Europium-151, properties
 RL: PRP (Properties)
 (151Eu Moessbauer spectra of quaternary barium/potassium europium titanium oxide **perovskites**)
- RE.CNT 31 THERE ARE 31 CITED REFERENCES AVAILABLE FOR THIS RECORD
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L8 ANSWER 48 OF 323 CAPLUS COPYRIGHT 2003 ACS on STN
AN 2001:72032 CAPLUS
DN 134:347503
TI High pressure synthesis and crystal structure of a new Ni(III)
perovskite: TlNiO₃
AU Kim, Seung-Joo; Demazeau, Gerard; Alonso, Jose A.; Choy, Jin-Ho
CS UPR-CNRS 9048, Institut de Chimie de la Matiere Condensee de Bordeaux
(ICMCB), Pessac, 33608, Fr.
SO Journal of Materials Chemistry (2001), 11(2), 487-492
CODEN: JMACEP; ISSN: 0959-9428
PB Royal Society of Chemistry
DT Journal
LA English
CC 78-2 (Inorganic Chemicals and Reactions)
Section cross-reference(s): 75, 77
AB A new Ni(III) compd., TlNiO₃ was prep'd. under high oxygen pressure. The
pressure and **temp.** domain required to prep. pure TlNiO₃
is very narrow (P .gtoreq. 7.5 GPa, 650 .ltoreq. T .ltoreq. 700.degree.).
TlNiO₃ crystallizes in the GdFeO₃-type **perovskite** structure with
the following lattice parameters: a 5.2549(1), b 5.3677(1) and c 7.5620(2)
.ANG.. The corresponding unit cell vol. of 213.3 .ANG.3 is somewhat
larger than the value expected from a plot of unit cell vol. vs. ionic
radius of the A cation for the analogous ANiO₃ series (A = rare earth
metal or yttrium). According to the structural anal., the twelve Tl-O
bonds in TlNiO₃ can be classified into three approx. groups (four short,
four medium and four long distances), while, in the other ANiO₃ compds.,
the A-O bonds are divided into two groups of eight short and four long
distances. The large unit cell vol. of TlNiO₃ is surely due to the
different coordination of the Tl(III) ion in the **perovskite**
lattice. The av. Ni-O-Ni bond angle (147.6.degree.) in TlNiO₃ is very
similar to that obsd. in YNiO₃ (147.3.degree.). However, the Neel temp.
(TN = 105 K) in antiferromagnetic TlNiO₃ is significantly lower than that
obsd. in YNiO₃ (TN = 145 K), which has a comparable Ni-O-Ni angle. The
magnetic properties could be explained by the competing effect of the
Tl(III)-O bond on the covalency of Ni(III)-O bond.
ST nickel thallium oxide **perovskite** prepn structure magnetic
property; crystal structure nickel thallium oxide **perovskite**
IT Crystal structure
Molecular structure
(of nickel(III) thallium(III) oxide **perovskite** (TlNiO₃))
IT Antiferromagnetic exchange
Exchange interaction
Magnetic susceptibility
Neel temperature
Thermal decomposition
(prepn., crystal structure, thermal decompr. and magnetic properties of
nickel(III) thallium(III) oxide **perovskite** (TlNiO₃))
IT Paramagnetism
(temp. independent; prepn., crystal structure, thermal decompr. and
magnetic properties of nickel(III) thallium(III) oxide
perovskite (TlNiO₃))
IT 12533-99-2P, Nickel thallium oxide (NiTlO₃)
RL: PRP (Properties); RCT (Reactant); SPN (Synthetic preparation); PREP
(Preparation); RACT (Reactant or reagent)
(prepn., crystal structure, thermal decompr. and magnetic properties of
nickel(III) thallium(III) oxide **perovskite** (TlNiO₃))
IT 1314-32-5, Thallium(III) oxide 12054-48-7, Nickel(II) hydroxide
RL: RCT (Reactant); RACT (Reactant or reagent)
(reactant for prepn. of nickel(III) thallium(III) oxide
perovskite (TlNiO₃))
RE.CNT 24 THERE ARE 24 CITED REFERENCES AVAILABLE FOR THIS RECORD
RE

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L8 ANSWER 72 OF 323 CAPLUS COPYRIGHT 2003 ACS on STN
AN 1999:806487 CAPLUS
DN 132:110086
TI High-pressure and -temperature synthesis and characterization of mixed valence perovskite oxides LaTi_{1-x}MgxO₃
AU Miao, J. -P.; Li, L. -P.; Song, Y. -B.; Xu, D. -P.; Lu, Z.; Su, W. -H.
CS Northeast Normal University and Department of Physics, Institute of Theoretical Physics, Jilin University, Changchun, Peop. Rep. China
SO Materials Chemistry and Physics (2000), 62(3), 226-229
CODEN: MCHPDR; ISSN: 0254-0584
PB Elsevier Science S.A.
DT Journal
LA English
CC 49-4 (Industrial Inorganic Chemicals)
AB Perovskite oxides LaTi_{1-x}MgxO₃ ($x = 0.25, 0.5$) were synthesized using high-pressure and -temp. method. LaTi_{0.75}Mg_{0.25}O₃ is a new compd. This new synthesis route has some advantages. XRD anal. showed that the $x = 0.25$ sample belongs to cubic perovskite-type structure and the $x = 0.5$ sample belongs to orthorhombic perovskite-type structure. EPR measurement indicated that Ti ions were in mixed valence state of +3 and +4. IR measurement indicated that the vibration frequency and width of B₀₆ octahedron stretching vibration absorption band decreases with the increasing of x . The valence state of Ti ions can be altered by high-pressure and -temp.
ST high pressure temp synthesis characterization mixed valence perovskite oxide
IT 12056-93-8P, Lanthanum magnesium titanium oxide (La₂MgTiO₆)
255916-99-5P, Lanthanum magnesium titanium oxide (LaMg_{0.25}Ti_{0.75}O₃)
RL: PNU (Preparation, unclassified); PREP (Preparation)
(high-pressure and -temp. synthesis and characterization of mixed valence perovskite oxides)
IT 12194-71-7P, Perovskite
RL: PNU (Preparation, unclassified); PREP (Preparation)
(mixed oxides of type of; high-pressure and -temp.
synthesis and characterization of mixed valence perovskite oxides)
RE.CNT 8 THERE ARE 8 CITED REFERENCES AVAILABLE FOR THIS RECORD
RE

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L8 ANSWER 81 OF 323 CAPLUS COPYRIGHT 2003 ACS on STN
AN 1999:364928 CAPLUS
DN 131:95831
TI Synthesis, structure and dielectric properties of Na₂SnTeO₆
AU Park, J.-H.; Woodward, P. M.; Parise, J. B.; Lubomirsky, I.; Stafsudd, O.
CS CHiPR, Dept. of Chemistry SUNY, Stony Brook, NY, 11794-2100, USA
SO Materials Research Society Symposium Proceedings (1999), 547(Solid-State
Chemistry of Inorganic Materials II), 139-144
CODEN: MRSPDH; ISSN: 0272-9172
PB Materials Research Society
DT Journal
LA English
CC 76-9 (Electric Phenomena)
Section cross-reference(s): 75
AB A new **perovskite** was recovered from the high **pressure**
-high **temp.** treatment of the .alpha.-TlSbO₃ form of Na₂SnTeO₆ at
7 GPa and 950.degree. for 30 min. Synchrotron x-ray powder diffraction
data show the space group is P21/n with a = 5.40361(5), b = 5.46152(5), c
= 7.69288(7) .ANG., and .beta. = 90.034(3).degree.. By using disk samples
of both polymorphs, the dielec. properties were measured as a function of
temp. At ambient conditions, the **perovskite** form has a more
than 1.5 fold enhancement in dielec. const. compared to the .alpha.-TlSbO₃
form while the molar volume and the mol. polarizability decrease.
ST sodium tin tellurate dielec property; crystal structure sodium tin
tellurate; synthesis sodium tin tellurate
IT Crystal structure
Dielectric constant
(synthesis, structure, and dielec. properties of Na₂SnTeO₆)
IT 229469-04-9
RL: FMU (Formation, unclassified); PRP (Properties); FORM (Formation,
nonpreparative)
(synthesis, structure, and dielec. properties of Na₂SnTeO₆)
RE.CNT 23 THERE ARE 23 CITED REFERENCES AVAILABLE FOR THIS RECORD
RE
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(22) Woodward, P; Acta Crystallogr 1997, VB53, P32 CAPLUS
(23) Woodward, P; submitted 1998

L8 ANSWER 149 OF 323 CAPLUS COPYRIGHT 2003 ACS on STN
AN 1996:663675 CAPLUS
DN 126:41762
TI High-pressure synthesis and thermal decomposition of LaCuO₃
AU Karppinen, M.; Yamauchi, H.; Ito, T.; Suematsu, H.; Fukunaga, O.
CS Laboratory of Inorganic and Analytical Chemistry, Helsinki University of
Technology, Espoo, FIN-02150, Finland
SO Materials Science & Engineering, B: Solid-State Materials for Advanced
Technology (1996), B41(1), 59-62
CODEN: MSBTEK; ISSN: 0921-5107
PB Elsevier
DT Journal
LA English
CC 78-2 (Inorganic Chemicals and Reactions)
AB The high-pressure synthesis of the rhombohedral LaCuO₃ **perovskite**
was studied and optimized in terms of the applied **pressure** and
temp., the synthesis time and the amt. of the oxidizer.
Essentially single phase samples were obtained at temps. >1400.degree.
under an O pressure of 5 GPa. The high O pressures were generated using
excess amts. of KClO₄ as an external oxidizing agent. Upon heating under
ambient pressure the stoichiometric high-pressure form of LaCuO_{3-y} loses O
yielding various O-deficient phases, and finally decompns. into La₂CuO₄ and
CuO or Cu₂O at .apprx.800.degree..
ST lanthanum cuprate prepn thermolysis high pressure; thermal decompr
lanthanum cuprate oxygen argon
IT Thermal decomposition
(of lanthanum cuprate in oxygen or argon)
IT Oxidizing agents
(potassium perchlorate in high pressure synthesis of lanthanum cuprate)
IT 1312-81-8, Lanthanum oxide (La₂O₃) 1317-38-0, Cupric oxide, reactions
RL: RCT (Reactant); RACT (Reactant or reagent)
(for prepn. of lanthanum cuprate at high pressure)
IT 7447-40-7P, Potassium chloride (KCl), preparation
RL: BYP (Byproduct); PREP (Preparation)
(formation in high pressure synthesis of lanthanum cuprate)
IT 1317-39-1, Cuprous oxide, formation (nonpreparative) 12053-92-8, Copper
lanthanum oxide (CuLa₂O₄)
RL: FMU (Formation, unclassified); FORM (Formation, nonpreparative)
(formation in thermal decompr. of lanthanum cuprate)
IT 37249-72-2D, Copper lanthanum oxide (CuLaO₃), oxygen-deficient
RL: FMU (Formation, unclassified); RCT (Reactant); FORM (Formation,
nonpreparative); RACT (Reactant or reagent)
(formation in thermal decompr. of lanthanum cuprate La₂CuO₄)
IT 37249-72-2P, Copper lanthanum oxide (CuLaO₃)
RL: PRP (Properties); RCT (Reactant); SPN (Synthetic preparation); PREP
(Preparation); RACT (Reactant or reagent)
(high-pressure synthesis and thermal decompr. of)
IT 7778-74-7, Potassium perchlorate (KClO₄)
RL: NUU (Other use, unclassified); USES (Uses)
(oxidant in high pressure synthesis of lanthanum cuprate)
IT 7440-37-1, Argon, uses 7782-44-7, Oxygen, uses
RL: NUU (Other use, unclassified); USES (Uses)
(thermal decompr. of lanthanum cuprate in atm. of)

L8 ANSWER 269 OF 323 CAPIUS COPYRIGHT 2003 ACS on STN
AN 1986:601893 CAPIUS
DN 105:201893
TI Effects of high pressure on some lanthanide ruthenium oxide systems:
synthesis of the **perovskite** phase praseodymium ruthenium oxide
(PrRuO₃)
AU Greatrex, R.; Hu, G.; Munro, D. C.
CS Dep. Inorg. Struct. Chem., Univ. Leeds, Leeds, LS2 9JT, UK
SO Materials Research Bulletin (1986), 21(7), 797-802
CODEN: MRBUAC; ISSN: 0025-5408
DT Journal
LA English
CC 78-2 (Inorganic Chemicals and Reactions)
Section cross-reference(s): 75, 76
AB The **perovskite** PrRuO₃ was prep'd. at pressures >2 GPa.
Lanthanide Ru oxide pyrochlores were studied at $\text{.1toreq.} 5.5 \text{ GPa}$ and
 $>1400.\text{degree.}$ The Pr compd. shows partial transformation to the
perovskite phase indicating some initial O deficiency. The PrRuO₃
perovskite exhibits elec. semicond. with activation energy
.apprx.0.17 eV.
ST elec cond praseodymium ruthenium oxide; lanthanide ruthenium oxide;
praseodymium ruthenium oxide **perovskite**
IT Crystal structure
(of praseodymium ruthenium oxide)
IT Electric conductivity and conduction
(of praseodymium ruthenium oxide with **perovskite**-type
structure)
IT Pyrochlore-type crystals
(rare earth ruthenium oxide, high-pressure effects on)
IT 105110-05-2P
RL: SPN (Synthetic preparation); PREP (Preparation)
(prepn. of, with **perovskite**-type structure at high pressures)
IT 37210-42-7P 61029-60-5P
RL: SPN (Synthetic preparation); PREP (Preparation)
(prepn. of, with pyrochlore-type structure at high pressures)
IT 61029-59-2P
RL: SPN (Synthetic preparation); PREP (Preparation)
(prepn. of, with pyrochlore-type structure with conversion to
perovskite-type compd. at high **pressures** and
temps.)
IT 12036-10-1
RL: RCT (Reactant); RACT (Reactant or reagent)
(reaction of, with rare earth metal oxides and ruthenium at high
pressures and **temps.**)
IT 1313-97-9 12037-29-5 12060-58-1
RL: RCT (Reactant); RACT (Reactant or reagent)
(reaction of, with ruthenium and ruthenium dioxide at high
pressure and **temp.**)
IT 7440-18-8, reactions
RL: RCT (Reactant); RACT (Reactant or reagent)
(reaction of, with ruthenium dioxide and rare earth metal oxides at
high **pressures** and **temps.**)

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L8 ANSWER 308 OF 323 CAPLUS COPYRIGHT 2003 ACS on STN
AN 1973:532377 CAPLUS
DN 79:132377
TI High-pressure synthesis of **perovskite** type rare earth indium oxide (RInO₃) (R=europium, gadolinium, and dysprosium)
AU Sawamoto, Hiroshi
CS Dep. Earth Sci., Nagoya Univ., Nagoya, Japan
SO Japanese Journal of Applied Physics (1973), 12(9), 1432-8
CODEN: JJAPA5; ISSN: 0021-4922
DT Journal
LA English
CC 78-4 (Inorganic Chemicals and Reactions)
Section cross-reference(s): 68, 70, 72
AB The stability regions of RInO₃ (R = Eu, Gd, and Dy) with **perovskite** structure were studied from 700-1200.degree. at pressures up to 90 kbar with the aid of the tetrahedral press. The stable pressure (kbar) and temp. (.degree.C) of **perovskite** structure were detd. as P > 9 + 0.07T for EuInO₃, P > 20 + 0.07T for GdInO₃, and P > 146 - 0.08T for DyInO₃. At normal pressure, rare earth indium garnet R₃(R_{1-x}In_{1+x})In₃O₁₂ was not synthesized, even though garnet solid solns. were synthesized at high pressure. Temp. variations of magnetic susceptibility for GdInO₃ and DyInO₃ follow Curie-Weiss law at >150.degree.K, with the effective Bohr magnetons of 5.83 .+- .0.17 for GdInO₃ and 9.03 .+- .0.05 for DyInO₃.
ST indium rare earth oxide; europium indium oxide; gadolinium indium oxide; dysprosium indium oxide; structure indium rare earth oxide; magnetism indium rare earth oxide
IT Magnetic moment
(of rare earth indium oxides)
IT Crystal structure
(of rare earth indium oxides and indium oxide-rare earth oxide solid solns.)
IT Dysprosium oxide (Dy₂O₃), solid soln. with indium oxide
Europium oxide (Eu₂O₃), solid soln. with indium oxide
Gadolinium oxide (Gd₂O₃), solid soln. with indium oxide
Indium oxide (In₂O₃), solid solns. with rare earth oxides
RL: PRP (Properties)
(crystal structure of)
IT 12292-63-6 12292-69-2 12292-84-1
RL: PRP (Properties)
(stability of, **pressure** and **temp.** effects on)

L8 ANSWER 309 OF 323 CAPLUS COPYRIGHT 2003 ACS on STN
AN 1973:458813 CAPLUS
DN 79:58813
TI Lead magnesium oxides prepared at high temperature and pressure
IN Hayashi, Hiroshi; Nishiyama, Goro; Nakayama, Noboru; Hasegawa, Kanemitsu;
Machida, Michihide
PA Agency of Industrial Sciences and Technology
SO Jpn. Kokai Tokkyo Koho, 2 pp.
CODEN: JKXXAF
DT Patent
LA Japanese
NCL 15P0
CC 70-1 (Crystallization and Crystal Structure)
Section cross-reference(s): 71
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 48005692	B4	19730124	JP 1971-39529	19710605
	JP 50029715		19750000	JP	

AB The crystal structure of Mg plumbate hydrate prep'd. from an aq. soln. of an alkali metal plumbate, and an aq. soln. of a Mg salt, such as MgCl₂ or Mg(NO₃)₂ were altered by simultaneous heating and pressing (1000.degree. and 50 kbars) for 20-60 min in a graphite tube to give MgPbO₃. Materials with **perovskite** (ABO₃) structure, such as BaTiO₃, SrTiO₃, CaTiO₃, CdTiO₃, and BaSnO₃ are useful as high-frequency dielec. and ferroelec. materials. X-ray diffraction data are given.
ST lead magnesium oxide prep'n; plumbate magnesium prep'n
IT X-ray
(diffraction of, by lead magnesium oxide)
IT 50864-25-OP
RL: SPN (Synthetic preparation); PREP (Preparation)
(prep'n. of, at high **pressure** and **temp.**)

Please forward to examiner. 466178

Hargrave, Sheena

From: STIC-ILL
Sent: Friday, September 26, 2003 11:44 AM
To: Hargrave, Sheena
Subject: FW: order

-----Original Message-----

From: Bos, Steven
Sent: Friday, September 26, 2003 11:42 AM
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Subject: order

****CA 68:71834****

DeVries et al
High pressure synthesis ..
J. of the American Ceramic Soc.
1968
51(2)
72-5

09/931312

Steven Bos
Art Unit 1754
308-2537

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Complete
faxed

L8 ANSWER 319 OF 323 CAPIUS COPYRIGHT 2003 ACS on STN
AN 1968:71834 CAPIUS
DN 68:71834
TI High-pressure synthesis of lead chromate (IV)
AU DeVries, Robert C.; Roth, Walter Lester
CS Gen. Elec. Co., Schenectady, NY, USA
SO Journal of the American Ceramic Society (1968), 51(2), 72-5
CODEN: JACTAW; ISSN: 0002-7820
DT Journal
LA English
CC 57 (Ceramics)
AB A new compd. with the compn. PbCrO₃, with Cr in the valence state of 4, was synthesized at high pressures above a **pressure-temp** line extending from .apprx.50 kilobars at 750.degree. to 60 kilobars at 1450.degree.. PbCrO₃ can be quenched and retained at 1 atm. but decomp. on heating >275.degree. at the same pressure. PbCrO₃ is considered to be an equil. phase at high pressures because it was synthesized from mixts. of PbO:CrO₂ as well as from several other mixts. of compds. in the Pb-Cr-O system. The new phase has the cubic **perovskite** structure and is the only known compd. with Cr⁴⁺ in an octahedral site. PbCrO₃ crystallizes primarily as black cubes which are often twinned on (111).
ST PRESSURE HIGH PB CHROMATES 4; CHROMATES 4 PB; HIGH PRESSURE PB CHROMATES 4; LEAD CHROMATES 4
IT 12218-29-0P
RL: SPN (Synthetic preparation); PREP (Preparation) (prepn. of)